## SILICON IN ORGANIC SYNTHESIS. 16. A SHORT SYNTHESIS OF $(\pm)-\alpha$ -VETISPIRENE<sup>1</sup>

Tu-Hsin Yan and Leo A. Paquette\*

Evans Chemical Laboratories, The Ohio State University, Columbus, Ohio 43210

Summary:  $\alpha$ -Vetispirene, a prototypical [4.5]spirobicyclic sesquiterpene, has been prepared in an efficient five-step reaction sequence beginning with  $\alpha$ -trimethylsilyl cyclopropanecarboxaldehyde. A new desiliconative alkylation is featured.

Although advantage has previously been taken of cyclopropane rings as building blocks in molecular construction, recognition of the special merits of silylcyclopropanes has gone virtually unrecognized. In the course of studies designed to develop the unique advantages offered by this class of reagents in synthesis, we initiated an investigation aimed at the stereoselective construction of quaternary carbon centers such as those which occur in the many known spirovetivane-type sesquiterpenes. The challenge posed by stereocontrolled spiroannulation has attracted considerable interest. Reported here is a particularly expedient synthesis of (±)-vetispirene (12) which highlights a new desiliconation-alkylation sequence of potentially wide applicability.

Our approach begins with the bifunctional reagent  $\frac{1}{2}$  which is readily available in 72% overall yield by modified Simmons-Smith cyclopropanation of 2-(trimethylsilyl)-2-propen-1-ol followed by oxidation with activated manganese dioxide. Coupling of  $\frac{1}{2}$  with 2,6-dimethylcyclohexenone ( $\frac{2}{2}$ ) in the presence of the Ti(0) reagent prepared by reduction of TiCl<sub>3</sub> provided the silylated diene  $\frac{3}{2}$  in 50-60% yield. Although the 300 MHz <sup>1</sup>H NMR spectrum of  $\frac{3}{2}$  (in C<sub>6</sub>D<sub>6</sub>) clearly revealed it to be a single stereoisomer, an unequivocal distinction between the two structural possibilities has not been made. Nor is it ultimately relevant.

Subjection of  $\frac{3}{2}$  to pyrolysis in a quartz chip-packed tube (30 cm long) at 30-40 torr (N<sub>2</sub> as carrier gas) expectedly required relatively high temperatures to achieve the vinylcyclopropane rearrangement. Under these conditions, smooth efficient bond reorganization occurred to yield

a 4:1 mixture of the spirocyclic vinylsilanes  $\frac{5}{2}$  and  $\frac{6}{2}$ . The dominance of  $\frac{5}{2}$  signaled preferential recombination of biradical  $\frac{4}{2}$  from that surface of the six-membered ring which is less sterically shielded.

The stage was now set for introduction of the remaining three carbon atoms. To this end,  $\frac{3}{2}$  was heated with anhydrous tetra-n-butylammonium fluoride and acetone in tetrahydrofuran solution at the reflux temperature for 10 h. These conditions serve to generate pentadienyl anion  $\frac{7}{7}$  which presumably experiences entirely regionselective alkylation at the cyclopropyl carbon atom in order to avoid the development of methylenecyclopropane character. Isomerically pure alcohol  $\frac{8}{7}$  was isolated in > 90% yield. Following conversion to methyl ether  $\frac{9}{7}$ , thermal rearrangement was effected at 435-440°C as before to produce in quantitative yield a mixture of  $\frac{10}{7}$  and  $\frac{11}{7}$  (1:5 ratio). The heightened stereoselectivity of this reaction, coupled with its efficiency and the ease with which  $\frac{11}{7}$  affords  $(\frac{1}{7})$ - $\alpha$ -vetispirene  $(\frac{12}{7}, \frac{100\%}{7})^{15}$  upon exposure to p-toluenesulfonic acid in benzene for 25-30 min at 5-20°C are particularly attractive. The overall yield for the five-step conversion of  $\frac{1}{7}$  into  $\frac{12}{7}$ , which can be executed on milligram and multigram scale as desired, was  $\frac{38\%}{7}$ .

Since difficulties are frequently encountered in the preparation of carbonyl activated cyclopropyl carbanions, the conversion of 3 to 8 represents a useful alternative to those synthetic
strategies which might normally require reactive intermediates of this type. The scope and
limitations of this C-C bond-forming process are presently under intensive investigation and
will be reported in due course.

Acknowledgment. This work was supported by the National Institutes of Health (Grants AI-11490 and GM-28468).

## References and Notes

- (1) Part 15. Paquette, L. A.; Daniels, R. G. Organometallics, in press.
- (2) Paquette, L. A. <u>Isr. J. Chem.</u> 1981, <u>21</u>, 128.
- (3) (a) Paquette, L. A.; Horn, K. A.; Wells, G. J. <u>Tetrahedron Lett.</u> 1982, 259; (b) Paquette, L. A.; Wells, G. J.; Horn, K. A.; Yan, T.-H. <u>Ibid.</u> 1982, 263.
- (4) Marshall, J. A.; Brady, S. F.; Anderson, N. H. <u>Fortschr. Chem. Org. Naturst.</u> 1974, 31, 283.
  - (5) For a representative selection of recent reports in this area, see: (a) Johnson, A.

- P.; Vajs, V. J. Chem. Soc. Chem. Commun. 1979, 817; (b) Murai, A.; Sato, S.; Masamune, T. Tetrahedron Lett. 1981, 1033; (c) Eilerman, R. G.; Willis, B. J. J. Chem. Soc. Chem. Commun. 1981, 30; (d) Ruppert, J. F.; White, J. D. J. Am. Chem. Soc. 1981, 103, 1808; (e) White, J. D.; Ruppert, J. F.; Avery, M. A.; Torii, S.; Nokami, J. Ibid. 1981, 103, 1813; (f) Burke, S. D.; Murtiashaw, C. W.; Dike, M. S.; Strickland, S. M. S.; Saunders, J. O. J. Org. Chem. 1981, 46, 2400; (g) Ficini, J.; Revial, G.; Genet, J. P. Tetrahedron Lett. 1981, 629, 633; (h) Iwata, C.; Ida, Y.; Miyashita, K.; Nakamischi, T.; Yamada, M. Chem. Ind. 1982, 165; (i) Subrahamanian, K. P.; Reusch, W. Tetrahedron Lett. 1978, 3789; (k) Baldwin S. W.; Fredericks, J. E. Ibid. 1982, 1235.
- (6) For previous syntheses of (±)-vetispirene, see: (a) Yamada, K.; Aoki, K.; Nagase, H.; Hayakawa, Y.; Hirata, Y. Tetrahedron Lett. 1973, 4967; (b) Dauben, W. G.; Hart, D. J. J. Am. Chem. Soc. 1975, 97, 1622; (c) Caine, D.; Boucugnani, A. A.; Chao, S. T.; Dawson, J. B.; Ing-walson, P. F. J. Org. Chem. 1976, 41, 1539; (d) Ibuka, T.; Hayashi, K.; Minakata, H.; Ito, Y. Inubushi, Y. Can. J. Chem. 1979, 57, 1579.
- (7) All new compounds exhibited compatible infrared, proton magnetic resonance, and mass spectroctopic data. In addition, the elemental composition of all key intermediates has been substantiated by combustion analysis.
  - (8) Sawada, S.; Inoue, Y. <u>Bull. Chem. Soc. Japan</u> 1969, 42, 2669.
  - (9) Chan, T. H.; Mychajlowskij, W.; Ong, B. S.; Harpp, D. N. <u>J. Org. Chem.</u> 1978, 43, 1526.
  - (10) Ohkata, K.; Paquette, L. A. J. Am. Chem. Soc. 1980, 102, 1082.
  - (11) Trost, B. M.; Salzmann, T. N.; Hiroi, K. J. Am. Chem. Soc. 1976, 98, 4887.
- (12) McMurry, J. E.; Krepski, L. R.; Fleming, M. P.; Kees, K. L. <u>J. Org. Chem.</u> 1978, 43, 3255; McMurry, J. E.; Krepski, L. R. <u>Ibid.</u> 1976, 41, 3929.
  - (13) Piers, E.; Lau, C. K. Synth. Commun. 1977, 7, 495.
- (14) The procedure employed is an adaptation of Sakurai's method for allyl anion generation [Hosomi, A.; Shirahata, A.; Sakurai, H. <u>Tetrahedron Lett.</u> 1978, 3043].
- (15) The IR and <sup>1</sup>H NMR spectra of our product proved identical to those supplied to us by Professor Drury Caine whom we thank.
- (16) E.g., cyclopropanecarboxylate esters: (a) Pinnick, H. W.; Chang, Y. H.; Foster, S. C.; Govindan, M. J. Org. Chem. 1980, 45, 4504; (b) Reissig, H.-U.; Böhm, I. J. Am. Chem. Soc. 1982, 104, 1735; (c) Kai, Y.; Knochel, P.; Kwiatkowski, S.; Dunitz, J. D.; Oth, J. F. M.; Seebach, D. Helv. Chim. Acta 1982, 65, 137. Compare lithium α-lithiocyclopropane carboxylates: (a) Ainsworth, C.; Kuo, Y. N. J. Organomet. Chem. 1973, 46, 73; (b) Warner, P. M.; Le, D. J. Org. Chem. 1982, 47, 893.

## (Received in USA 6 May 1982)